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Section 13.11: Gas Stoichiometry Using $PV=nRT$ we can define the volume occupied by one mole of ANY gas at STP: $V=(nRT)/P$
 $V=(1 \text{ mol})(0.08206 \text{ Latm/ Kmol})(273\text{K})/1 \text{ atm}$ $V=22.4 \text{ L}$ The value of 22.4 L/mol (@STP) is called the molar volume. It is true for ANY ideal gas (or real gas behaving ideally as they do under STP conditions). 13-35

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Gases Chapter 13 - Folsom Cordova Unified School District

13.1 Gases and Their Properties 487 For an ideal gas (in which the particles occupy no volume and experience no attractions or repulsions), gas pressure and volume are inversely proportional. This means that if the temperature and the number of gas particles are constant and if the volume

Chapter 13 Gases - An Introduction to Chemistry

Chapter 13 - States of Matter - 13.1 The Nature of Gases - 13.1 Lesson Check - Page 424: 3 Answer The kinetic theory assumes that gases are tiny particles with insignificant volume and are always in rapid, constant, random motion.

Chapter 13 - States of Matter - 13.1 The Nature of Gases

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13 1 Chapter 13 States Of Matter 13 1 The Nature Of Gases ...

13 Answer Key. 13-1 Explore, page 13-1 "Review and Reinforcement, additional volumes of the valve, the higher gas pressure on the liquid proportional to Because gases are very compressible, the particles must be 44 Chapter 13.

Chapter 13 study guide gases section 13.1...

Chapter 13 Gases. STUDY. PLAY. Barometer. is a device that measures atmospheric pressure. mm Hg. is millimeters of mercury; a unit of measurement for pressure, also called torr; $760 \text{ Hg} = 760 \text{ torr} = 101,325 \text{ Pa} = 1 \text{ standard atmosphere (atm)}$. Torr. is another name for millimeters of mercury (mm Hg).

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Chapter 13 Gases. STUDY. PLAY. The Kinetic Molecular Theorem(KMT) A theory that explains the behavior of gases and the observed properties of gases at the molecular level. Normal conditions. Low pressure $>10\text{atm}$ and high temp well above condensing point of gas(0C) and 0p for most room temp gases; under normal conditions real gases behave as ...

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Chapter 13 Gases Part 1. STUDY. PLAY. 5 assumptions of the kinetic molecular theory of gases. 1. gases consist of many particles that are spread apart relative to their size. 2. collisions between particles and the walls of the container are elastic.

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13.1 The Gas Laws. •Absolute zero is zero on the Kelvin scale. • Charles's law states that the volume of a given amount of gas is

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directly proportional to its kelvin temperature at constant pressure. SECTION. 13.1 The Gas Laws Charles's Law. (cont.) Gay-Lussac's Law.

Chemistry: Matter and Change

Chapter 13: Gases. Hard Chem test! Woah! This can be hard because it's like definitions and stuff so it's hard to do the exact one. STUDY. ... (postulate of the kinetic molecular theory of gases) The particles that make up gases are so small, compared with the distances between them, that the volume is assumed to be...

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Section 13.6 Exercise 27.4 L of oxygen gas at 25.0°C and 1.30 atm, and 8.50 L of helium gas at 25.0°C and 2.00 atm were pumped into a tank with a volume of 5.81 L at 25°C. • Calculate the new partial pressure of oxygen. 6.13 atm • Calculate the

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new partial pressure of helium. 2.93 atm • Calculate the new total pressure of both gases. 9 ...

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Chapter 13 - Gases 197 Reread the Study Sheets in this chapter and decide whether you will use them or some variation on them to complete the tasks they describe. Study Sheet 13.1: Using the Ideal Gas Equation Sample Study Sheet 13.2: Using the Combined Gas Law Equation Sample Study Sheet 13.3: Equation Stoichiometry Problems

Chapter 13 - Gases

Start studying Chapter 13 The Gas Laws Test. Learn vocabulary, terms, and more with flashcards, games, and other study tools.

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Chapter 13 Gas Laws 13-1 Objectives C4.5a Provide macroscopic

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examples, atomic and molecular explanations, and mathematical representations (graphs and equations) for the pressure-volume relationship in gases.

Chapter 13 Gas Laws Notes - Chapter 13 Gas Laws 13-1 ...

What pressure would this sample of gas exert in a 1.50-L container at the same temperature? 6. A gas has a volume of 5.0 L at a certain pressure. How must the pressure be changed to double the volume of the gas at constant temperature? a. The pressure must be doubled. ... CHAPTER 13 Practice Test - Gases

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484 Chapter 13 Gases 13.1 Gases and Their Properties If you want to understand how gases behave—such as why fresh air rushes into your lungs when certain chest muscles contract or how gases in a car's engine move the pistons and power the

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car—you need a clear mental image of the model chemists

Chapter 13 Gases - An Introduction to Chemistry

Section 13.1 continued CHAPTER 13 STUDY GUIDE FOR CONTENT MASTERY Vacuum Atmospheric pressure Pressure exerted by mercury column 760 mm Name Date Class Study Guide for Content Mastery Chemistry: Matter and Change Chapter 13 73 States of Matter Section 13.1 Gases In your textbook, read about the kinetic-molecular theory. Complete each statement. 1.

13 STUDY GUIDE FOR CONTENT MASTERY 13 STUDY GUIDE FOR ...

Section 13-1 Section 13.1 The Gas Laws State the relationships among pressure, temperature, and volume of a constant amount of gas. scientific law: describes a relationship in nature that is...

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386 Chapter 13 Section 13.1 (continued) Gas Pressure
TEACHERTEACHER Demo Elastic Collisions Purpose Students will differentiate elastic collisions from perfectly elastic collisions.
Materials Newtonian cradle (a device in which small steel balls are suspended by thin nylon tethers to horizontal wooden sticks) Procedure Explain that in an elastic

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