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the steady state processes since involving differential equations which can be solve by integration. These calculations with their applications in many chemical engineering fields (mass transfer, heat transfer, chemical kinetics,…etc.) will be given in "Applied Mathematics in Chemical Engineering" within 3rd year of study. Chapter 7

Basic Principles and Calculations in Chemical Engineering

Welcome to Basic Principles and Calculations in Chemical Engineering. Several tools exist in the book in addition to the basic text to aid you in learning its subject matter. We hope you will take full advantage of these resources. Learning Aids 1.

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Prof. Manolito E Bambase Jr. Department of Chemical Engineering. University of the Philippines Los Baños SLIDE 14 Example 11-3. Combustion of Methane (CH 4) Simplifying the equations (1) C: 1 = Px 1 (2) H: 4 = Px 4 (3) O: 4.34 = 2Px 2 + 2Px 1 + Px 4 (4) N: 16.34 = 2Px 3 (5) x: x 1 + x 2 + x 3 + x 4 = 1

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h h Area = ½ bh a2 = b2 + c2 – 2bc·cos∠A b2 = a2 + c2 – 2ac·cos∠B c2 = a2 + b2 – 2ab·cos∠C h b a c A B C Perimeter = 2a + 2b Ellipse Area = ab 2b n = number of sides f s Rectangle Circle Parallelogram Area = bh h b h Pyramid A = area of base Solid Geometry Sphere Volume = s Volume r3 Surface Area = 4 r2 r h w d Rectangular Prism Volume = wdh

Engineering Formula Sheet - madison-lake.k12.oh.us
It can be determined by dividing the molecular weight by the number of hydrogen atoms or hydroxyl ions (or their equivalent) supplied or required by the molecule in a given reaction. B. (redox reaction) the molecular weight in grams divided by the change in oxidation state. Ion product of water (K.

ChemE
Prof. Manolito E Bambase Jr. Department of Chemical Engineering. University of the Philippines Los Baños SLIDE 5 Dehydrogenation of Ethane Total Balance: Input = Output Molecular Species Balance: C 2H 6: Input – Consumed = Output C 2H 4: Generated = Output H 2: Generated = Output Atomic (Elemental) Species Balance: C-Balance: Input = Output

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By algebra, h = 0.53cG1/2/[(cμ/k)3/4(D/μ)1/2], and by substitution, h = 0.53c gβρL2 D T 1/2 1/2 (cμ/ k)3/4(D/μ)1/2 = 0.53c gβρL2 D 1/4 T 1/4 (cμ/

k)3/4(D/μ)1/2Thus, h = 0.53(1.0)[(4.18 × 108)(0.0004)(60)2(1/12)]1/4 T 1/4 [1.0(0.72)/0.395]3/4[(1/12)/0.72]1/2 = 83.57 T 1/4where h is in Btus per hour per square foot per degree Fahrenheit, and the temperature difference isin degrees Fahrenheit.4.

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