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Awareness: How will students be able to recognize the activities that you are doing? How will they know what the lesson objectives are and what they are supposed to be doing during the lessons?

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The Activity Series Use with Chapter 9,
Section 9.2 1. For each of the following
pairs of elements, underline the one that
would replace the other element in a
compound. a. calcium, tin e. iron, copper
b. bromine, fluorine f. iodine, chlorine c.

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aluminum, potassium g. silver, lead d.
zinc, sodium 2.

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Class STUDY GUIDE In your textbook, read about the activity series for metals and halogens. Then examine each of the following pairs of potential reactants. Use Figure 9.13 in your textbook to help you decide whether or not a reaction would occur. If a reaction occurs, write the balanced equation.

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Solutions Manual Chemistry: Matter and Change • Chapter 9 143 CHAPTER 9 SOLUTIONS MANUAL Problem-Solving

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Lab page 294 Halogen Atomic Radius
(ppm) Ionic Energy (kJ/mol)

Electronegativity Fluorine 72 1681 3.98

Chlorine 100 1251 3.16 Bromine 114

1140 2.96 Iodine 133 1008 2.66 Astatine

140 920 2.2 Properties of Halogens 1.

Chemical Reactions **Chemical Reactions**

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Worksheet Balancing Word Equations
Chapter 10 (Remember the following are
diatomic: H₂, N₂, O₂, F₂, Cl₂, Br₂, I₂)
2) The coefficients should add up to the

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number at the end that is in parenthesis.
1. carbon + oxygen carbon dioxide (0) 2.
copper + silver nitrate copper(II) nitrate
+ silver (4) 3.

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with Chapter 8, Section 8.3 TEACHING
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Model and Molecular Shape 30 Use with
Chapter 8, Section 8.4 1.

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29 Lewis Structures W TT ...

Use with Chapter 3, Section 3.4 S (-QS 2.

What is the difference between a mixture and a pure substance? Q 3.

What is the difference between a compound and an element? 4. Can a compound be a heterogeneous mixture?

Explain your answer by referring to the ... Teaching Transparency Worksheets

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Chemistry: Matter and Change Chapter 3

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9. Is the outer electron configuration of the O ion after the reaction a very stable one? How can you tell? CD 10. What is the electron configuration of the Ca ion?

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Chemistry: Matter and Change Chapter 7

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